2010 Scheme

QP CODE: 202006 **Reg. No: ……..**

Second Year B.Pharm Degree Supplementary Examinations September 2021

Pharmaceutical Analysis I

Time: 3 Hours **Total Marks: 100**

- ***Answer all questions to the point neatly and legibly • Do not leave any blank pages between answers • Indicate the question number correctly for the answer in the margin space***
- ***Answer all parts of a single question together • Leave sufficient space between answers***
- ***Write equations wherever necessary.***

Essay **(3x10=30)**

1. Explain the neutralization curves of different acid-base titrations and suggest suitable indicators for detection of end point.
2. Types of solvents used in non-aqueous titrations. Explain levelling and differentiating effect of non-aqueous solvent. Give the preparation and standardization and 0.1m perchloric acid.
3. Explain argentometric titration and its types. Explain briefly Mohr's method.

Short notes **(14x5=70)**

4. Explain the principle involved in Law of Mass Action.
5. Explain the chromophore theory of pH indicators with examples.
6. Redox indicators with example.
7. Give an account on the preparation and standardization of sodium thiosulphate solution.
8. Methods of expressing concentration of solutions.
9. Types of errors and steps to be taken to minimize errors in an analytical procedure.
10. Explain the various steps involved in gravimetric analysis.
11. Explain the different types of complexometric titrations by using various titrants with suitable examples.
12. Advantages of titrations involving cerric ammonium sulphate
13. Discuss about oxygen flask combustion method.
14. Explain about the calibration of pipette.
15. Preparation of standard solution of potassium permanganate. Why freshly prepared solutions of potassium permanganate is heated prior to its use in titrations.
16. pM indicators.
17. Accuracy and precision.

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